

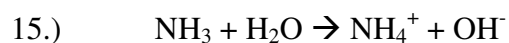
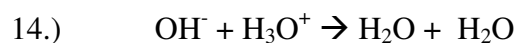
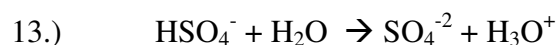
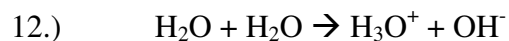
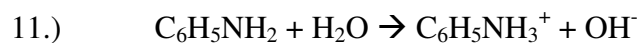
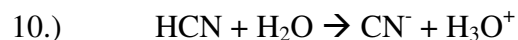
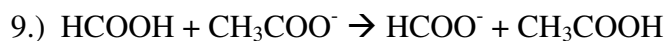
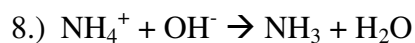
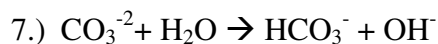
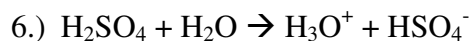
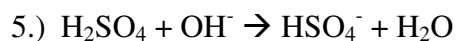
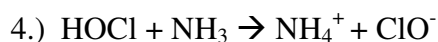
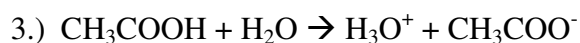
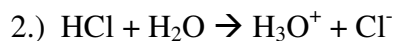
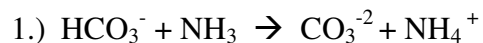
Worksheet – Bronsted-Lowry Acids and Bases

Name _____

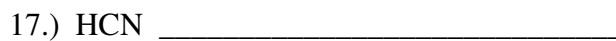
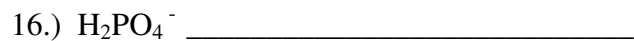
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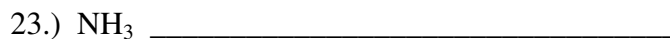
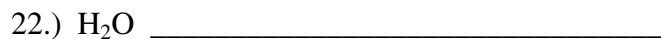
Identify the conjugation acid-base pairs in the following reactions. An acid donates a proton to become a conjugate base. A base accepts a proton to form a conjugate acid.



Write the formulas for the conjugate bases formed by each of the following acids.



Write the formulas for the conjugate acids formed by each of the following bases.



24.) After an acid has given up its proton, it is capable of getting back that proton and acting as a base. A conjugate base is what is left after an acid gives up a proton. The stronger the acid, the weaker the conjugate base. The weaker the acid, the stronger the conjugate base. Complete the table below.

Acid	Base	Equation
H_2SO_4	HSO_4^-	$\text{H}_2\text{SO}_4 \rightarrow \text{H}^+ + \text{HSO}_4^-$
H_3PO_4		
	F^-	
	NO_3^-	
H_2PO_4^-		
H_2O		
	SO_4^{2-}	
HPO_4^{2-}		
NH_4^+		
	H_2O	